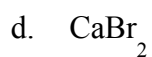
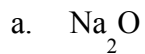


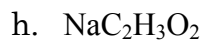
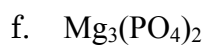
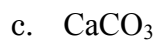
Unit 9 – Exercise 3 - Ionic Compounds

1. Describe the properties of ionic compounds and the evidence that supports those properties.

2. Give the name of the following binary ionic compounds.



3. Give the name of the following polyatomic ionic compounds.



4. Write the formula for the following binary ionic compounds.
- a. lithium bromide
 - b. sodium iodide
 - c. calcium sulfide
 - d. cesium oxide
 - e. beryllium iodide
 - f. lithium acetate
 - g. rubidium sulfate
 - h. strontium nitrate
 - f. barium chloride
 - g. aluminum fluoride
 - h. potassium oxide
5. Write the formula for the following polyatomic ionic compounds.
- a. ammonium sulfide
 - b. calcium hydroxide
 - c. potassium nitrate
 - d. sodium carbonate
 - e. aluminum phosphate