

1. Alloy	A mixture of two or more metals	24. Dilute	Having a relatively low concentration of solute in a mixture
2. Amorphous Solid	A solid made up of particles that are not arranged in a regular pattern	25. Ductile	Can be drawn into wires
3. Atom	Smallest particle of an element	26. Element	A pure substance made of only one kind of atom
4. Axis	A horizontal or vertical line drawn on a graph to define points of reference.	27. Evaporation	Vaporization that takes place only on the surface of a liquid
5. Best-Fit Line	The line that most closely approximates the data in a scatter plot	28. Freezing	The change of state from a liquid to a solid
6. Boiling	Vaporization that occurs on and below the surface of a liquid	29. Gas	A state of matter with no definite shape or volume
7. Boiling Point	The temperature at which a substance changes from a liquid to a gas	30. Graph	Represents numerical information spatially
8. Brittle	Easily broken	31. Heating Curve	A plot of temperature versus time for a substance where energy is added at a constant rate
9. Chemical Bond	The force that holds two atoms together	32. Heterogeneous	Having parts that are unrelated or completely different
10. Chemical Change	A change in matter that produces one or more new substances	33. Homogeneous	Of the same kind; similar; uniform
11. Chemical Energy	Energy stored in chemical bonds	34. Hypothesis	A proposed, scientifically testable explanation for an observed phenomenon.
12. Chemical Properties	Characteristic that cannot be observed without altering the substance	35. Independent Variable	The experimental factor that is manipulated; the variable whose effect is being studied.
13. Colloid	A mixture containing small, undissolved particles that do not settle out	36. Liquid	A state of matter that has no definite shape but has a definite volume
14. Compound	A substance made up of atoms of two or more different elements joined by chemical bonds	37. Luster	The way a mineral reflects light (shininess)
15. Concentrated	Having a relatively large amount of substance present in a unit amount of mixture	38. Malleable	Capable of being shaped
16. Conductivity	The ability of an object to transfer heat or electricity to another object	39. Mass	The amount of matter in an object
17. Controlled Experiment	An experiment in which only one variable is manipulated at a time.	40. Matter	Anything that has mass and takes up space
18. Crystalline Solid	A solid whose atoms, ions, or molecules are arranged in an orderly, geometric structure	41. Melting	The change in state from a solid to a liquid
19. Data	Facts and statistics collected together for reference or analysis	42. Melting Point	The temperature at which a substance changes from a solid to a liquid
20. Density	The degree of compactness of a substance. Equal to a substance's mass divided by its volume	43. Metal	Elements that are good conductors of heat and electric current
21. Dependent Variable	The measurable effect, outcome, or response in which the research is interested.	44. Metalloid	An element that has properties of both metals and nonmetals
22. Deposition	The change of state from a gas directly to a solid	45. Mixture	A combination of two or more substances that are not chemically combined
23. Diatomic Molecule	A molecule consisting of two atoms	46. Molecule	Smallest unit of most compounds, made of two or more atoms chemically bonded together
		47. Nonmetal	An element that lacks most of the properties of a metal
		48. Physical Changes	A change in a substance that does not change its identity

49. Physical Properties	Characteristics that can be observed without changing the identity of the substance
50. Plasma	A high temperature state of matter in which an atom is stripped of its electrons
51. Pure Substance	Matter that always has exactly the same composition
52. Saturated	A solution that has reached its maximum solubility
53. Semiconductor	A substance that can conduct electricity under some conditions
54. Slope	The steepness of a line on a graph. Equal to the line's rise divided by its run.
55. Solid	Definite shape and volume
56. Solubility	A measure of how much solute can dissolve in a given solvent at a given temperature
57. Solubility Curve	Graph representing the solubility of a saturated solution at a variety of temperatures
58. Solute	A substance that is dissolved in a solution
59. Solution	A homogeneous mixture of two or more substances
60. Solvent	Substance doing the dissolving; largest portion of the solution
61. State of Matter	A physical property that describes matter as a solid, liquid, or gas
62. Sublimation	A change directly from the solid to the gaseous state without becoming liquid
63. Supersaturated	Contains more dissolved solute than a saturated solution at the same temperature
64. Suspension	A mixture in which particles can be seen and easily separated by settling or filtration
65. Unsaturated	A solution that is capable of dissolving more solute at that temperature
66. Vaporization	The change of state from a liquid to a gas
67. Viscosity	A liquid's resistance to flowing
68. Volume	The amount of space an object takes up
69. Weight	A measure of the force of gravity on an object
70. Y-Intercept	The y-coordinate of a point where a graph crosses the y-axis