

1. Actinides	The 14 elements with atomic numbers from 90 to 103 (the bottom of the bottom two rows of the periodic table)	25. Molecular Compound	Atoms in a compound bonded with covalent bonds, where the atoms share valence electrons
2. Alkali Metal	An element in Group 1 of the periodic table	26. Neutron	A subatomic particle that has no charge and that is found in the nucleus of an atom
3. Alkaline Earth Metal	An element in Group 2 of the periodic table	27. Noble Gas	An element in Group 18 of the periodic table
4. Anion	A negatively charged ion	28. Nomenclature	Naming system
5. Atomic Number	The number of protons in the nucleus of an atom	29. Nonpolar	Having an even distribution of charge
6. Bohr Model of the Atom	Arranged the electrons in circular orbits around the nucleus	30. Nucleus	Center of an atom
7. Cation	A positively charged ion	31. Nuclide	A general term for a specific isotope of an element
8. Cloud Model of the Atom	Electrons occupy orbitals around nucleus.	32. Orbital	A region in an atom where there is a high probability of finding electrons
9. Covalent Bond	A chemical bond that involves sharing a pair of electrons between atoms in a molecule	33. Oxidation Number	Tells you how many electrons an atom has gained, lost, or shared to become stable
10. Crystal	An orderly, three-dimensional pattern of ions or atoms in a solid	34. Periodic Table	A chart of the elements showing the repeating pattern of their properties
11. Dalton Model of the Atom	Solid sphere; different for every substance	35. Periods	The horizontal rows of elements on the periodic table
12. Dot Diagram	A diagram to represent electrons in the outer energy level of an atom	36. Polar	Having an uneven distribution of charge
13. Double Bond	A chemical bond formed when atoms share two pairs of electrons	37. Polyatomic Ion	An ion made of two or more atoms; acts as a single unit
14. Electron	A tiny, negatively charged particle that moves around the nucleus of an atom.	38. Proton	Positively charged subatomic particle
15. Electron Cloud	A region around the nucleus of an atom where electrons are likely to be found	39. Rare Earth metals	Lanthanides and actinides (the bottom two rows of the periodic table)
16. Group	A column on the periodic table	40. Reactivity	How readily a substance combines chemically with other substances.
17. Halogen	An element in Group 17 of the periodic table	41. Rutherford Model of the Atom	Electrons orbit the nucleus
18. Ion	A charged atom	42. Sea of Electrons	Valence electrons of metal atoms, flows and transmits electrons
19. Ionic Bond	Formed when one or more electrons are transferred from one atom to another	43. Thomson Model of the Atom	"Plum Pudding", electrons embedded in sphere
20. Ionic Compound	A compound that consists of positive and negative ions	44. Transition Metal	One of the elements in Groups 3 through 12 of the periodic table.
21. Isotope	Atoms of the same element that have different numbers of neutrons	45. Triple Bond	A chemical bond formed when atoms share three pairs of electrons
22. Lanthanides	The 14 elements with atomic numbers from 58 to 71 (the top of the bottom two rows)	46. Valence Electrons	Electrons on the outermost energy level of an atom
23. Mass Number	The sum of the number of neutrons and protons in an atomic nucleus		
24. Metallic Bond	An attraction between a positive metal ion and the electrons surrounding it		