**Key Terms:**

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| --- | --- | --- | --- |
| * Anion
 | * Cation
 | * Covalent Bond
 | * Crystal
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| * Double Bond
 | * Electricity
 | * Ion
 | * Ionic Bond
 |
| * Mineral
 | * Nonpolar
 | * Polar
 | * Polyatomic Ion
 |
| * Triple Bond
 | * Molecular Compound
 |

**Test Material:**

* Be able to identify metals and nonmetals on the periodic table.
* Be able to determine the charge of an ion based on the group of the element on the periodic table.
* Understand how an element forms an ion and how to determine whether its charge will be positive or negative.
* Know the groups of the most reactive metals and the most reactive nonmetals
* Be able to identify whether a compound is ionic or molecular from its name and/or formula.
* From a compound’s structure, be able to identify which would have the lowest/highest melting point.
* Given a formula, be able to name the compounds.
* Given a name, be able to write the formula for the compound
* Be able to compare/contrast ionic and molecular compounds (at least 2 differences and 2 similarities).
* Understand the differences in the naming systems for ionic compounds and molecular compounds. Be able to compare and contrast them.
* Be able to describe what happens if an ionic crystal is struck with a hammer and why.
* Understand how soap works and be able to explain it in terms of polarity.
* Be able to compare and contrast ionic and molecular crystals.
* Be able to use a reference table including the periodic table and a table of ions.